Course: Introductory course in general inorganic chemistry designed for the student with little or no previous science training.

Description: Lecture: M – R 11:00 – 12:15 AM in Chem 190

Sections: Discussions 601 – 604

Instructor: Dr. Thomas Sorensen Phone: 229-4012

Office Hours: MWF 8:30 – 9:20 AM in Chem 109

Course web-site, see D2L: d2l.uwm.edu

Prerequisites: A math placement score of 30 or a grade of C or better in Math 105. Not open for credit to students who have credit in Chem 102(ER), 111(ER), or 117(ER).

In order for you to be successful in this course you will be required to identify and set-up problems and perform algebraic manipulations.

Course Materials: See D2L for additional details and options.

<table>
<thead>
<tr>
<th>Component</th>
<th>Estimated Cost</th>
<th>Comments</th>
</tr>
</thead>
<tbody>
<tr>
<td>ALEKS 360</td>
<td>$70 – $130</td>
<td>Required, code includes an ebook</td>
</tr>
<tr>
<td>Non-Programmable Calculator</td>
<td>$10 – $50</td>
<td>Required for exams and quizzes</td>
</tr>
</tbody>
</table>

**No graphing/programmable calculators will be allowed on quizzes or exams!**

Bauer, Birk, and Marks, 4th Ed. $25 – $215 Recommended (printed version of the ebook)

Policies: UWM: You must follow the policies and procedures outlined in the current Schedule of Classes. See: http://www.uwm.edu/Dept/SecU/SyllabusLink.pdf

Department of Chemistry and Biochemistry: You are expected to fully understand these policies including the limits placed on the maximum amount of material that can be missed, excused or otherwise, and still pass the course.

Drop, Section Change: These are done on PAWS. Make sure to follow all the rules established by UWM and the Department of Chemistry and Biochemistry.

Incomplete: An Incomplete can be given only to a student who has been doing satisfactory (C) work but who is unable to continue attending the course for a reason judged valid. The request for an Incomplete must be accompanied by documentation.

Academic Dishonesty: Cheating on an examination or other graded material will result in a grade of zero as a minimum consequence. Failure in the course and referral to the Dean may also occur. In short, academic dishonesty in any form will not be tolerated.

Attendance: It will be a significant advantage for you to attend every lecture. You are responsible for all material presented in lecture. If you miss a lecture, you are responsible for obtaining the lecture material.

Discussion: Bring your completed Lecture Exercises to each discussion as well as your calculator. There may be a quiz/worksheet given in each of your scheduled Discussion meetings. These components will be combined and recorded four times during the semester. The average of these four grades constitutes your Discussion grade.

Homework: You will need to establish an ALEKS account and complete your Initial Assessment right away (see D2L for details). Most Objectives are due on Sunday at 11:59 PM, except the last one which is due the last day of class (the night before Study Day) but check ALEKS for official due dates.

Your homework grade is based on the average of all your ALEKS objectives and the total number of topics you have mastered/learned as of the last day of the semester (the day before the Study Day).

Lecture: There will be a quiz/attendance given/taken at least once during most lectures. The average of these constitutes your Lecture grade.
Early/Make-Up/Late Work: There are no early, make-up, or late exams, homework, or quizzes. For a scheduled absence (e.g. University athletics, music, etc.), the instructor must be notified at least 24 hours prior to the absence. If an exam or quiz is missed for a non-medical reason not approved beforehand, a grade of zero (0) will be given. For medical absences, a written letter signed by your physician is required. For an excused absence the next quiz or exam will count double.

Assessment: Your course grade will be determined from the following elements:

<table>
<thead>
<tr>
<th>Element</th>
<th>Points</th>
</tr>
</thead>
<tbody>
<tr>
<td>Discussion</td>
<td>100 pts</td>
</tr>
<tr>
<td>Homework</td>
<td>150 pts</td>
</tr>
<tr>
<td>Lecture</td>
<td>50 pts</td>
</tr>
<tr>
<td>Hour Exam I</td>
<td>100 pts</td>
</tr>
<tr>
<td>Hour Exam II</td>
<td>100 pts</td>
</tr>
<tr>
<td>Hour Exam III</td>
<td>100 pts</td>
</tr>
<tr>
<td>Redemption Exam</td>
<td>Extra Credit</td>
</tr>
<tr>
<td>Final Exam</td>
<td>200 pts</td>
</tr>
</tbody>
</table>

Assessment: Your course grade will be determined from the following elements:

<table>
<thead>
<tr>
<th>Week of</th>
<th>Lecture</th>
<th>ALEKS^x</th>
</tr>
</thead>
<tbody>
<tr>
<td>June 26 MT  lec01 – lec03</td>
<td>Initial Assessment</td>
<td></td>
</tr>
<tr>
<td>WR       lec02 – lec03</td>
<td>Obj1</td>
<td></td>
</tr>
<tr>
<td>July 3 MT    lec04 – lec05</td>
<td></td>
<td></td>
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<tr>
<td>WR       lec06</td>
<td>Obj2^b</td>
<td></td>
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<tr>
<td>Exam I, Friday, 7-7-17, 11:00 – 12:15 PM through lec06</td>
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<tr>
<td>10 MT       lec07 – lec09</td>
<td></td>
<td></td>
</tr>
<tr>
<td>WR       lec10 – lec12</td>
<td>Obj3</td>
<td></td>
</tr>
<tr>
<td>17 MT       lec13 – lec15</td>
<td></td>
<td></td>
</tr>
<tr>
<td>WR       lec16 – lec17</td>
<td>Obj4^b</td>
<td></td>
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<tr>
<td>Exam II, Friday, 7-21-17, 11:00 – 12:15 PM through lec17</td>
<td></td>
<td></td>
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<tr>
<td>24 MT       lec18 – lec19</td>
<td></td>
<td></td>
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<tr>
<td>WR       lec20 – lec22</td>
<td>Obj5</td>
<td></td>
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<tr>
<td>31 MT       lec23 – lec25</td>
<td></td>
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<tr>
<td>WR       lec26</td>
<td>Obj6</td>
<td></td>
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<tr>
<td>Exam III, Friday, 8-4-17, 11:00 – 12:15 PM through mod26</td>
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<tr>
<td>Aug 7 MT     lec27 – lec29</td>
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<tr>
<td>WR       lec30 – lec32</td>
<td>Obj7</td>
<td></td>
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<tr>
<td>14 MT       Review</td>
<td></td>
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<tr>
<td>Redemption Exam, Wednesday, 8-16-17, 11:00 – 12:15 PM Covers lectures and the text</td>
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<tr>
<td>Final Exam (covers lectures and the text)</td>
<td>Friday, 8-18-17, 11:00 – 12:15 PM</td>
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</table>

^x See ALEKS for exact due dates. Start assignments well before their due date.

^b An ALEKS assessment is required before you are allowed to start the next assignment.

You will find that your understanding of the material increases as you work more problems. Please get help when you need it. Sometimes 10 – 15 minutes spent in the TA or my office saves hours of confusion and frustration. I will attempt to remain on schedule as much as possible, but changes may be made with reasonable notice.

Some Advice: We will cover a large amount of material in this course. The amount of time you need to spend reading and solving problems is significant. You are expected to read the material before each lecture and should not expect all assigned material to be explicitly covered in class. Write down questions that you have and note things that you don’t understand; bring your questions to me or your TA for discussion. If
you are having trouble with the material, you must seek help fast or the rapid pace of the course will leave you far behind. The TAs and I are available and eager to help you.

**How to do well in the course:**

- Read the appropriate section(s) of the book before the material is presented in lecture — that is, come prepared for class.
- Attend lecture and discussion sessions — and take them seriously. Be punctual and take notes.
- Do problems as they are assigned — *don’t wait* until right before an exam when they may seem overwhelming.
- Form a *study group* with others in the class and work on homework questions together. Teaching one another is perhaps the best way to learn.
- Try to make your own connections between material presented during different lectures. Don’t just assume because we say that something is connected that you understand the connections.
- If you have problems, *see me and/or your TA* and find out the ways that we can help you.

**D2L Resources include:**

- Information on ALEKS and TopHat.
- PowerPoint slides (as PDF).
- Your grades.
Chemistry 100 Lectures for Chemical Science (Chem 100), Semester III, 16-17

Reading and associated problems from the textbook by lecture.

lec01: **Introduction to Chemistry**
Reading- Ch01: 1.1, 1.2 (p 13, 18, 19, 21, 23 – 27), and Ch02: 2.5 (p 75 – 78). Qualitative only, calculations will be done later.

lec02: **Atomic Theory — The Basics**
Reading- Ch02: 2.1 and 2.2.
Problems- Ch02: 3 – 42.

lec03: **Atoms and Ions**
Reading- Ch02: 2.3.

lec04: **Atomic Theory — Isotopes**
Reading- Ch02: 2.4.

lec05: **Light, Electrons, and the Bohr Model**
Reading- Ch07: 7.1 and 7.2 (ignore calculations for now).
Problems- Ch07: 1 – 36.

lec06: **Electron Configurations and the Periodic Table**
Reading- Ch07: 7.3 and 7.4.
Problems- Ch07: 37 – 70.

lec07: **Hund’s Rule, Valence and Core Electrons**
Reading- Ch07 7.5 and 7.6.
Problems- Ch07: 71 – 84.

lec08: **Lewis Dot Structures**
Reading- Ch03: 3.1, Ch08: 8.1 (p 304 – 306) and 8.3 (p 313 – 322).
Problems- Ch03: 5 – 7, Ch08: 3 – 12, and 45 – 64.

lec09: **Molecular Shape**
Reading- Ch08: 8.5.
Problems- Ch08: 83 – 94.

lec10: **Polarity**
Reading- Ch08: 8.1 (p 307 – 309) and 8.5 (p 333 – 335).
Problems- Ch08: 17 – 26, 107 – 112, 119, and 120.

lec11: **Intermolecular Forces**
Reading- Ch10: 10.2 (p 410 – 417).
Problems- Ch10 41 – 74.

lec12: **Ionic Compounds and Chemical Formulas**
Reading- Ch08: 8.2, Ch03: 3.2 and 3.3.
Problems- Ch08: 28, 35, 36, 40 – 42, and CH03: 11 – 38.
lec13: **Periodic Trends**
   Reading- Ch07: 7.7.
   Problems- Ch07: 85 – 94 and 102 – 108.

lec14: **Nomenclature**
   Reading- Ch03: 3.5.

lec15: **Algebraic Review and Scientific Notation**
   Reading- Math Tool Box 9.2 (p 379 and 380) and Math Tool Box 1.1 (p 35 and 36).
   Problems- Ch01: 3 – 8.

lec16: **Significant Figures**
   Reading- Math Tool Box 1.2 (p 37 – 41).
   Problems- Ch01: 9 – 14.

lec17: **Unit Conversions – Dimensional Analysis**
   Reading- Math Tool Box 1.3 (p 41 – 45).
   Problems- Ch01: 19, 20, 23, 24, 122, 128, and 131 – 133.

lec18: **Density and the Temperature Scales: Fahrenheit, Celsius, and Kelvin**
   Reading- Ch01: 1.2 (p 16 – 23).
   Problems- Ch01: 70 – 78 and 83 – 87.

lec19: **Formula Calculations – The Mole**
   Reading- Ch04: 4.2.
   Problems- Ch04: 9 – 30.

lec20: **Formula Calculations – The Mole, Part II**
   Reading- Ch04: 4.2.
   Problems Ch04: 36 – 48 and 53 – 60.

lec21: **Percent Composition and Empirical Formulas**
   Reading- Ch04: 4.1 and 4.3.
   Problems- Ch04: 3, 4, 8, 67, 68, 73, and 74.

lec22: **Empirical Formulas and Molecular Formulas**
   Reading- Ch04: 4.3 (p 148 and 149).
   Problems- Ch04: 82 – 85 and 87 – 89.

lec23: **Solutions**
   Reading- Ch04: 4.4.
   Problems- Ch04: 91 – 114.

lec24: **Chemical Reactions – General**
   Reading- Ch05: 5.1 – 5.3 and 5.4 (p 183 – 185).
   Problems- Ch05: 3 – 52.

lec25: **Chemical Reactions – Single Displacement**
   Reading- 5.4 (p 190 – 193).
   Problems- Ch05: 61 – 68.
mod26: Chemical Reactions – Double Displacement – Precipitation Reactions
   Reading- Ch05: 5.4 (p 193 – 196).

mod27: Mole and Mass Ratios
   Reading- Ch06: 6.1 – 6.3.

mod28: Formula Units – Molecules
   Reading- Ch06: 6.1 – 6.3.

mod29: Limiting Reactant, Theoretical Yield, and Percent Yield
   Reading- Ch06: 6.4 and 6.5.

mod30: Solutions and Stoichiometry
   Reading- Ideas from Ch04 and Ch06.
   Problems- Handout.

mod31: Gases – General Principles and Concepts
   Reading- Ch09: 9.1 – 9.3.
   Problems- Ch09: 1 – 36 and 81 – 93.

mod32: Gases – Ideal Gas Law and Stoichiometry
   Reading- Ch09: 9.1 – 9.3.
   Problems- Ch09: 37 – 80 and 115 – 118.